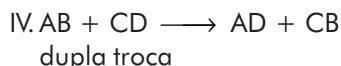
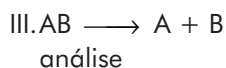
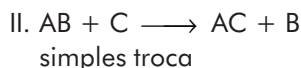
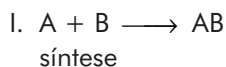


Testes

Aula 25

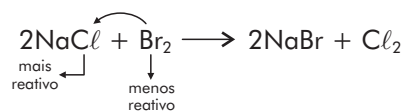
25.01) D



25.02) A

exotérmico \longrightarrow Libera calor.

25.03) C



Nessa equação, não ocorre propriedade química característica dos halogênicos.

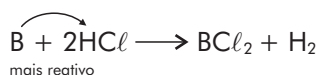
25.04) 42



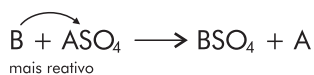
menos reativo
mais nobre
A reação não ocorreu.

25.05) 10

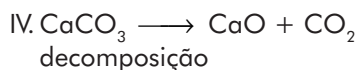
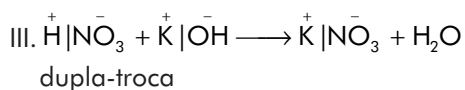
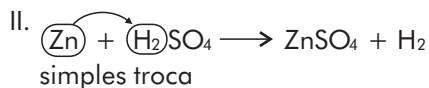
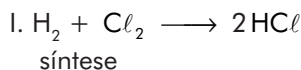
01. **Incorreta.**
02. **Correta.**



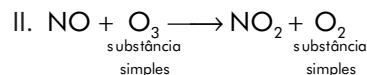
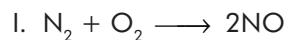
04. **Incorreta.**
08. **Correta.**



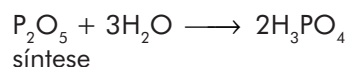
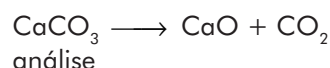
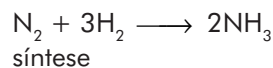
25.06) A



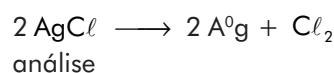
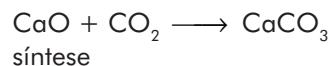
25.07) E



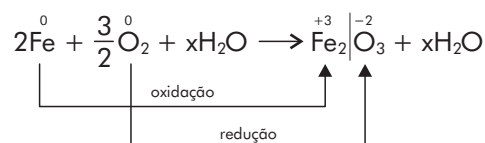
25.08) A



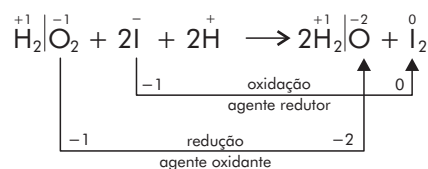
25.09) A



25.10) B

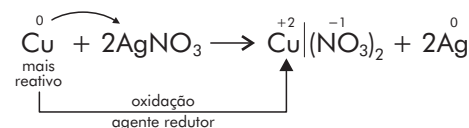


25.11) C

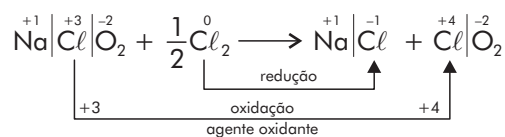


$\left\{ \begin{array}{l} \text{H}_2\text{O}_2 - \text{agente oxidante} \\ \text{I}^- - \text{agente redutor} \end{array} \right.$

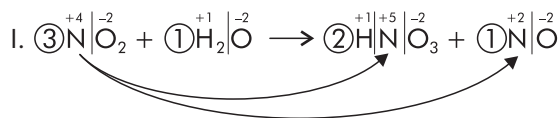
25.12) B



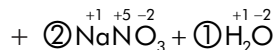
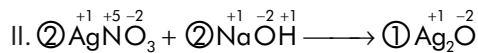
25.13) C



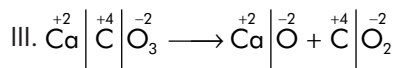
25.14) A



{ Muda o Nox
redox

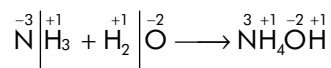


Não muda o Nox.



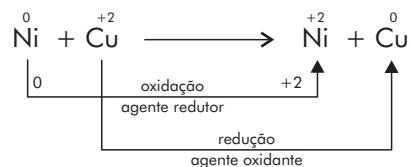
Não muda o Nox

25.15) C



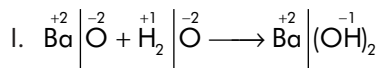
Não muda o Nox. (Não e Redox.)

25.16) B

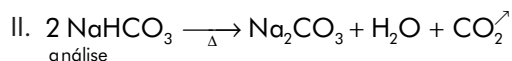


Ni { Sofre oxidação agente redutor.

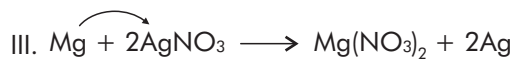
25.17) A



síntese

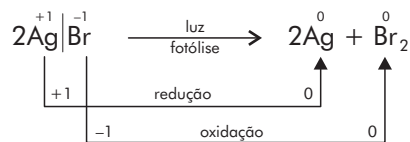


análise

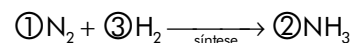


mais reativo
simples troca

25.18) C

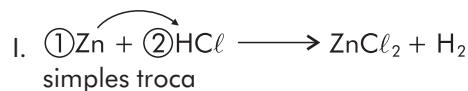


25.19) C

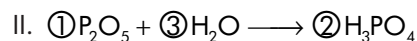


{ NH₄OH – hidróxido de amônio
NH₃ – amônia

25.20) E



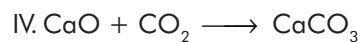
simples troca



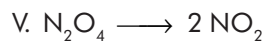
síntese



dupla-troca



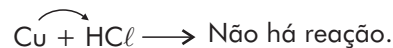
síntese



{ análise
decomposição

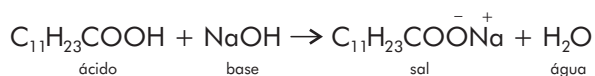
Aula 26

26.01) 19



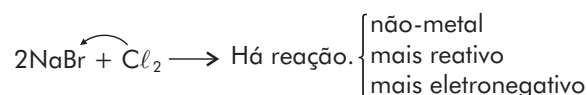
menos reativo
mais nobre

26.02) A

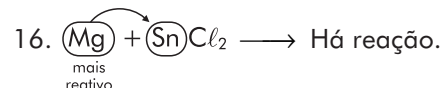
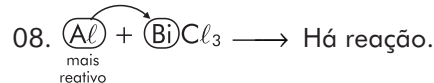
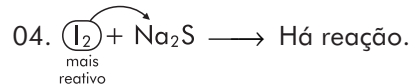
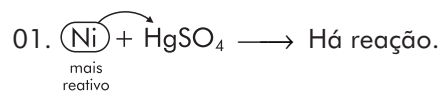


dupla-troca

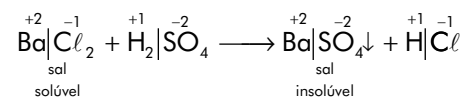
26.03) A



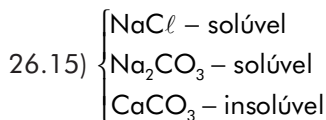
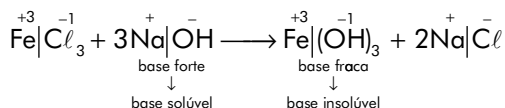
26.04) 29



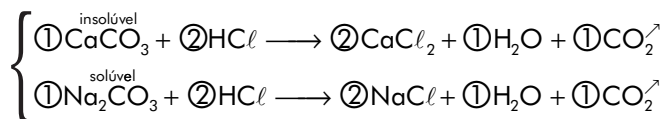
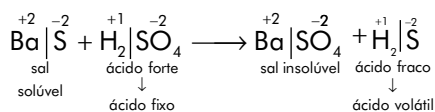
26.05) B



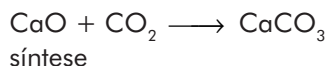
26.06) C



26.07) A



26.08) A

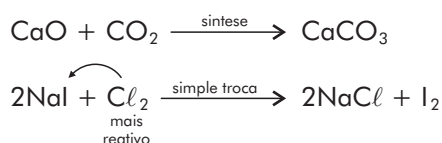


Colocando-se os sais em água, apenas o CaCO_3

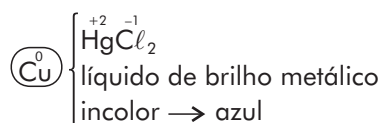
não se dissolverá. Adicionando-se HCl aos demais, somente o Na_2CO_3 produzirá efervescência.

cia.

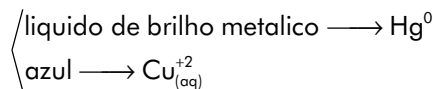
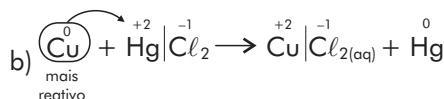
26.09) B



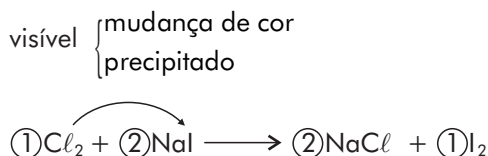
26.16) Discursivo



a) Redox (simple troca)

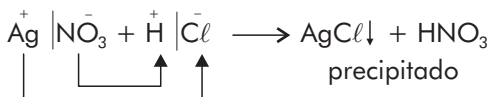


26.10) B

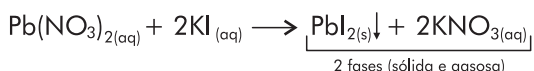


O BaSO_4 é insolúvel e, portanto, há diminuição da concentração iônica responsável pela condução elétrica. No instante em que todo o H_2SO_4 for consumido, a condutibilidade elétrica será praticamente nula. Com excesso de $\text{Ba}(\text{OH})_2$, haverá íons desse eletrólito na solução e a condutibilidade elétrica começa a aumentar.

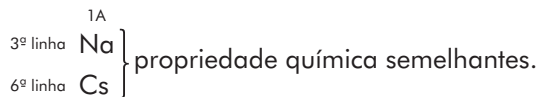
26.11) D



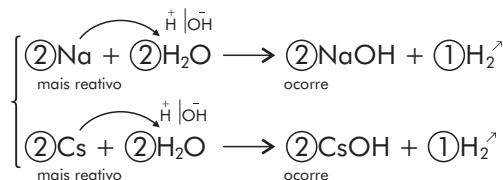
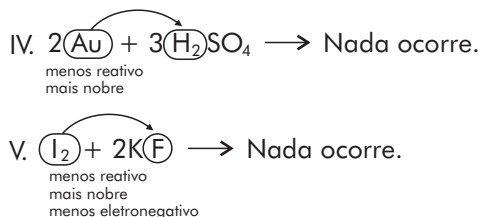
26.12) A



26.18) A



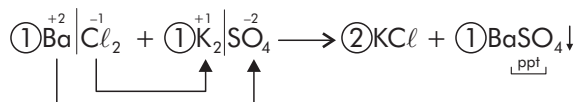
26.13) D



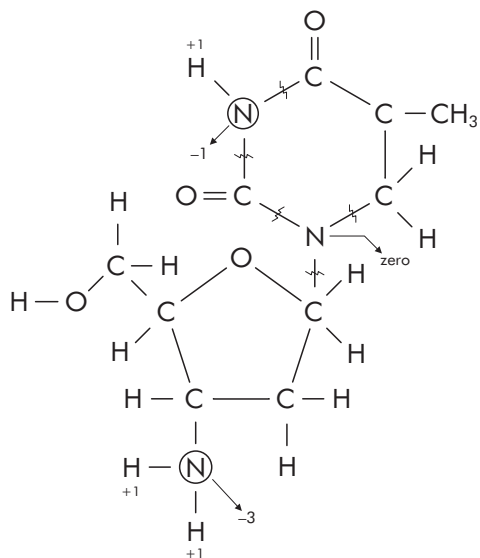
26.14) 31

- I. $\text{AB} + \text{CD} \longrightarrow \text{AC} + \text{BD}$
dupla-troca
- II. $\text{AB} + \text{C} \longrightarrow \text{AC} + \text{B}$
simple troca
- III. $\text{A} + \text{B} \longrightarrow \text{AB}$
síntese
- IV. $\text{CD} \longrightarrow \text{C} + \text{D}$
decomposição

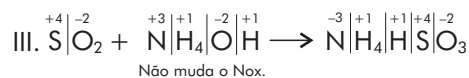
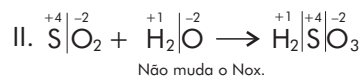
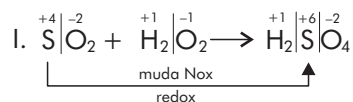
26.19) C



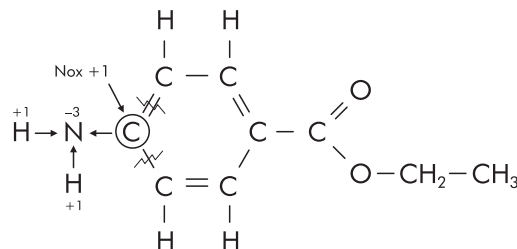
27.09) D



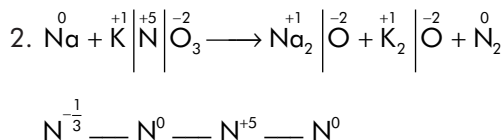
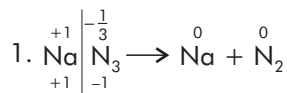
27.15) A



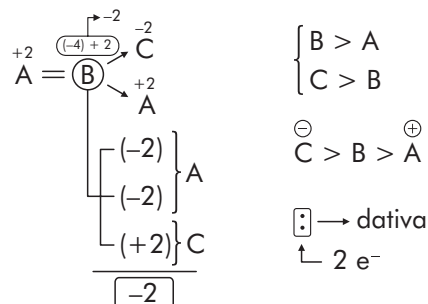
27.16) A



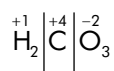
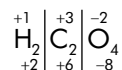
27.10) B



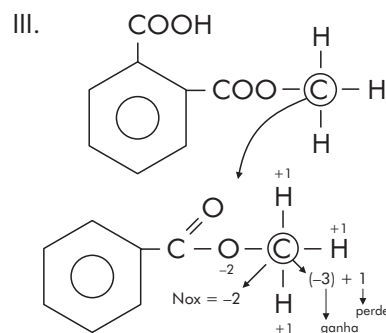
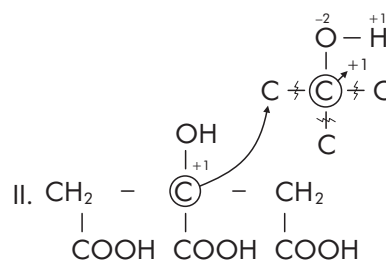
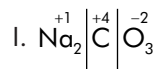
27.17) B



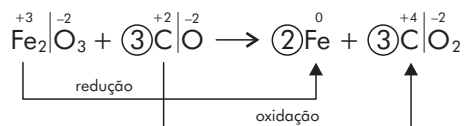
27.11) B



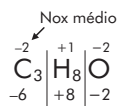
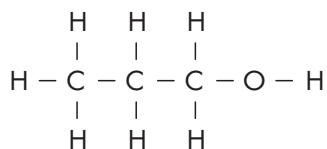
27.18) B



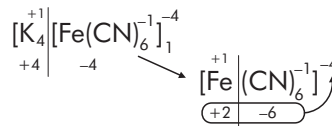
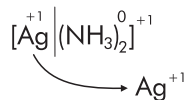
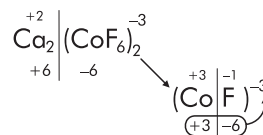
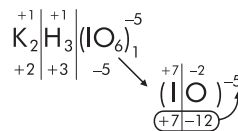
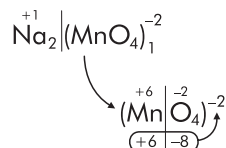
27.14) E



27.19) D

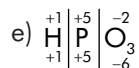
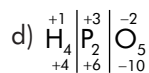
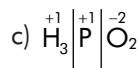
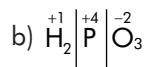
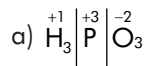


27.20) D



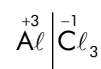
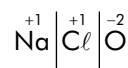
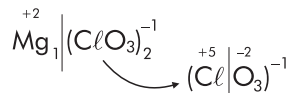
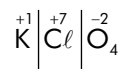
Aula 28

28.01) E

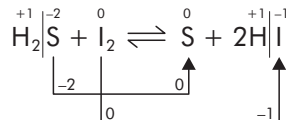


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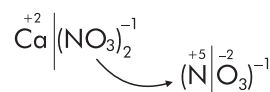
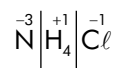
28.02) B



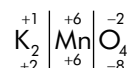
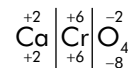
28.03) E



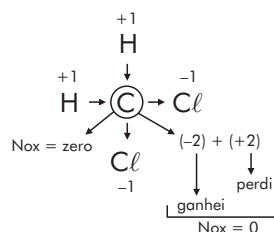
28.04) C



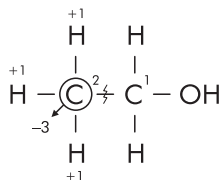
28.05) D



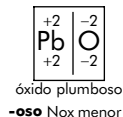
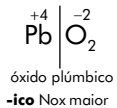
28.06) C



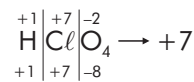
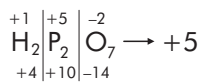
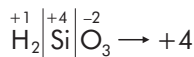
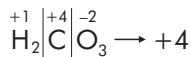
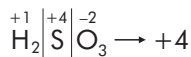
28.07) D



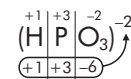
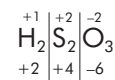
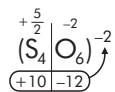
28.08) C



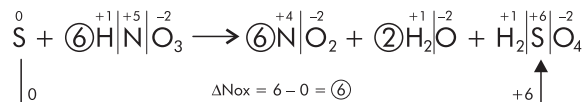
28.09) C



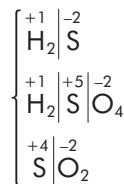
28.10) D



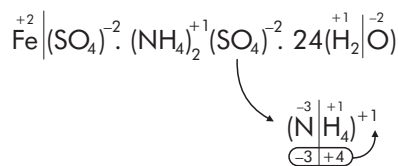
28.11) E



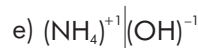
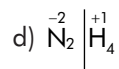
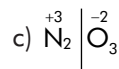
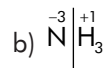
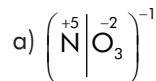
28.12) C



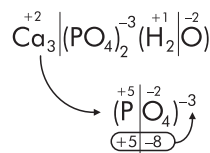
28.13) C



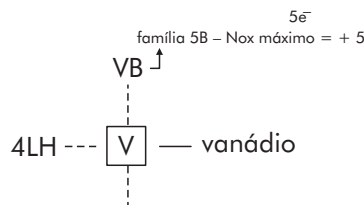
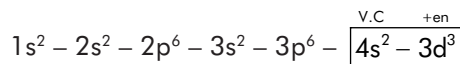
28.14) D



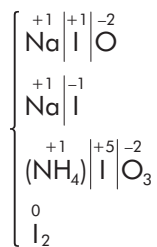
28.15) D



28.16) D

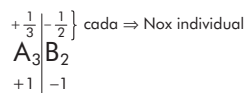


28.17) C

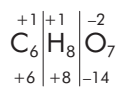


$$\cancel{+1} + (\cancel{-1}) + (+5) + 0 = 5$$

28.18) D



28.19) B



28.20) E

